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## Chemistry Study Guide

Remember that while this CORRECTLY COMPLETED study guide has a lot of valuable information that will be on the test, you are encouraged to also study notes in SI, handouts, worksheets, and labs in order to best prepare for the test.

1. A characteristic of matter that can be seen or measured WITHOUT CHEMICALLY CHANGING that matter is a \_\_\_\_\_ property.
2. A characteristic of matter that can only be observed when one substance changes into or produces something new is a \_\_\_\_\_ property.
3. A change in which the identity of the substance **does NOT** change is known as a \_\_\_\_\_ change.
4. A change in which one type of matter transforms into **another type or kind that may have very different properties** from the original is known as a \_\_\_\_\_ change.
5. The state of matter in which the molecular configuration is the most spread out because the atoms have the most energy and are moving fast is the \_\_\_\_\_ state.
6. The state of matter that is created when heat (more energy) is added to a solid is the \_\_\_\_\_ state.
7. The state of matter with the least energy in which the atoms are tightly packed together and vibrate slowly in place is the \_\_\_\_\_ state.
8. A pure substance in its simplest form (a single atom) that cannot be broken down into any other substance by chemical or physical means is a(n) \_\_\_\_\_.
9. A pure substance made of two or more elements chemically combined in a set ratio is a \_\_\_\_\_.
10. When one or more of the same element combine and that substance can exist alone and keep its original properties is a \_\_\_\_\_.

11. In the chart below, put a check mark in the appropriate category for the symbol/formula listed. **Remember, all compounds are molecules, but not all molecules are compounds.**

Symbol/Formula	Atom (Element)	Molecule	Compound
NaCl			
C			
MgSO <sub>4</sub>			
O <sub>2</sub>			

## BIOLOGY: Chapter 34 Assessment-Reptiles &amp; Birds

## True/False

Indicate whether the sentence or statement is true or false.

1. Both reptiles and birds lay amniotic eggs.
2. A reptile's high metabolism causes most reptiles to be endothermic.
3. Reptiles have dry, largely watertight skin and lay watertight eggs.
4. Reptiles must return to the water in order to reproduce.
5. The surviving reptiles include crocodiles and alligators, turtles, the tuatara, and snakes.
6. Feathers may be specialized for flight.
7. Down feathers cover the body of adult birds.
8. Birds have a three-chambered heart.
9. The shape of a bird's beak is a strong indicator of its food source.

## Multiple Choice

Identify the letter of the choice that best completes the statement or answers the question.

10. Reptiles have
  - a. internal fertilization.
  - b. endothermic metabolism.
  - c. concurrent flow.
  - d. cartilaginous skeletons.
11. Which of the following is a reptilian adaptation to living on land?
  - a. external fertilization
  - b. endothermic temperature regulation
  - c. respiration through gills
  - d. the amniotic egg
12. Unlike other reptiles, turtles and tortoises
  - a. live only in water.
  - b. are prehistoric.
  - c. are endangered.
  - d. do not have teeth.
13. Snakes are reptiles of the order
  - a. Squamata.
  - b. Chelonia.
  - c. Crocodylia.
  - d. Rhynchocephalia.
14. Birds retain many reptilian features, including
  - a. teeth.
  - b. a long bony tail.
  - c. scales on their feet and lower legs.
  - d. None of the above
15. Birds are different from reptiles in that they
  - a. are endothermic.
  - b. have feathers, rather than scales, covering their bodies.
  - c. have four-chambered hearts.
  - d. All of the above
16. The bones of birds
  - a. are composed primarily of keratin.
  - b. are solid.
  - c. are found sparingly throughout the body.
  - d. are thin and honeycomb
17. A bird's heart has

## Chapter 7, Worksheet 2

Name \_\_\_\_\_

## AP Chem, Periodic Properties of Elements

Period \_\_\_\_\_

1. As a general rule, the size of atoms increases as you move to the (left/right or up/down) on the periodic table. What part of this generalization is counter-intuitive?

2. Predict the atom in each of the following pairs with the larger radius.

- a) Be or Mg    b) C or N    c) Se or Cl  
d) Ba or Y    e) Mg or Sc    f) Te or I

3. As a general rule, metals that are (farther from/closer to) the boundary between metals and nonmetals have more "metallic character", and nonmetals that are (farther from/closer to) the boundary have more "nonmetallic character".

4. Predict the element with more metallic character.

- a) Cu or Sr    b) Al or Mg    c) Rb or K    d) Rb or Ba

5. Predict the element with more nonmetallic character.

- a) N or O    b) S or Se    c) As or F    d) B or Si

6. As a general rule, (metal/nonmetal) ions are larger than (metal/nonmetal) ions.

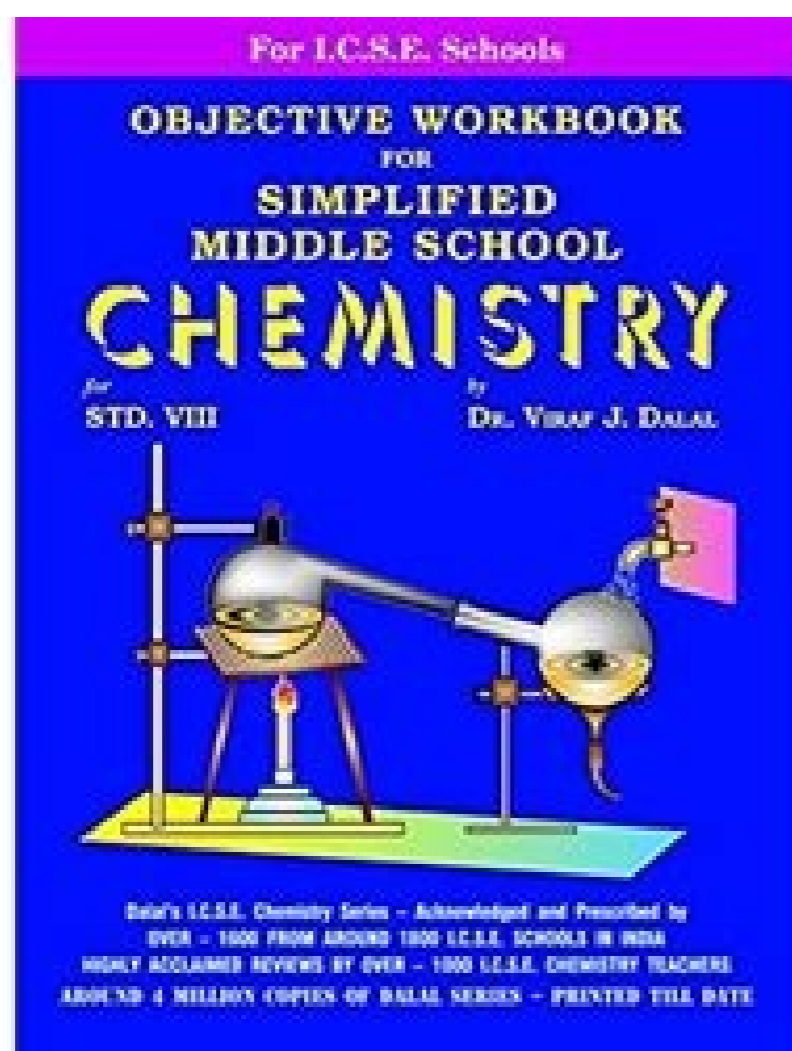
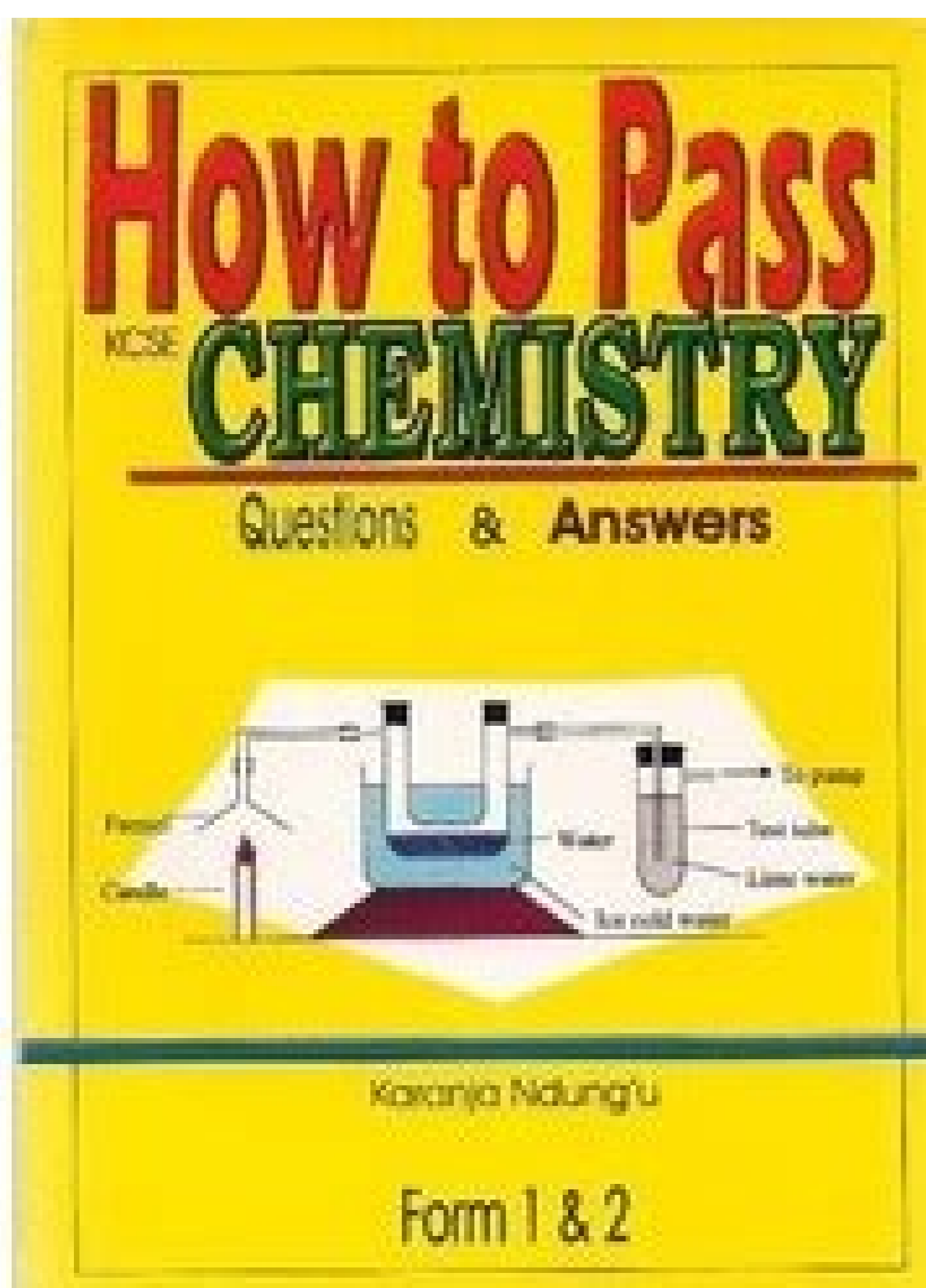
7. As a general rule, metal atoms are (larger/smaller) than their (+) ions, and nonmetal atoms are (larger/smaller) than their (-) ions.

8. In general, the size of other positively charged or negatively charged ions increases as you move to the (left/right or up/down) on the periodic table.

9. Predict the larger of each of these pairs of atoms or ions.

- a) Mg or Mg<sup>2+</sup>    b) Br or Br<sup>-</sup>    c) K or K<sup>+</sup>    d) N<sup>3-</sup> or Mg<sup>2+</sup>  
e) Na<sup>+</sup> or F<sup>-</sup>    f) Sr<sup>2+</sup> or Ba<sup>2+</sup>    g) Mg<sup>2+</sup> or Al<sup>3+</sup>    h) I or I<sup>-</sup>  
i) F or Cl<sup>-</sup>    j) Al<sup>3+</sup> or P<sup>3-</sup>

10. As a general rule, first ionization energy (the energy needed to remove the first electron from a neutral gaseous atom) increases as you move to the (left/right or up/down) on the periodic table.



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You've just added this product to the cart: IGCSE Chemistry Revision Answers Test 1 Experimental Techniques Test 3 Atoms, Elements and Compounds 1. (a), (b) and (d) 2. (b) 3. 1 mortar 2 thermometer 3 tripod stand 4 Bunsen burner 4. (a) 5. T, T, T, F 6. A Obtaining ethanol from a mixture of water and ethanol. B Separating the different colours in a sample of ink. C Obtaining sand from a mixture of sand and seawater. D Obtaining large copper sulphate crystals from an aqueous solution of copper sulphate. Test 2 Experimental Techniques 1. (c) 2. (b), (c), and (d) 3. (a) 4. (d) 5. (c) 6. Nucleus; electrons; electrons; neutrons 7. (a) 8. F, T, F Test 3 Atoms, Elements and Compounds 1. 1 The dark brown colour of the liquid is seen just above the dish. 2 A medium brown colour is seen close to the dish. 3 A light brown colour is seen some distance from the dish. 4 The same light brown colour is seen throughout the beaker. 2. (b) 3. (d) 4. 1. graphite; 2. diamond; 3. graphite; diamond; 4. diamond; 5. graphite; 6. diamond 5. -2 6. (c) 7. (c) Test 4 Atoms, Elements and Compounds Test 8 Substances and Stoichiometry 1. (c) 2. 17 3. (b) 4. (c) 5. T, F, F, T, T, 6. (c) 7. 2; calcium; 20 8. (b) 9. (b) Test 5 Atoms, Elements and Compounds Test 9 Substances and Stoichiometry 1. (c) 2. (b) 3. (c) 4. (c) 5. (c) 6. (a) 7. Radioactive; radioactive; nuclei; radiation; measuring; medical; treating; sterilising 8. 1 Hydrogen chloride gas 2 Water 3 Hydrogen chloride dissolved in water 1. (a) 2. (c) 3. (b) 4. (b) 5. (a), (c) and (f) 6. (a) 7. T, T, F, F, F, 8. (d) 9. (b), (c) and (e) Test 7 Substances and Stoichiometry 1. (c) 2. (a) 3. (c), (d) and (e) 4. (d) 5. F, F, T, F, T, T 6. (d) 7. (d) 8. (a) and (b) 9. T Test 11 Electricity and Chemistry 1. F, F, T, 2. (b) 3. (a) and (d) 4. Metals; hydrogen; non-metals; hydrogen 5. (c) 6. T, F, F, T 7. (b) 8. (b) and (c) Test 12 Electricity and Chemistry 1. F, T, F, 2. (d) 3. (b) 4. (b) and (d) 5. (b) and (c) Test 13 Electricity and Chemistry 1. Hydrogen 2. (a) 3. Gain; reduced 4. T, T, F, F, T 5. T, 6. A Impure copper B Pure copper C Copper sulphate solution 7. Negatively; ions; electrons Test 14 Electricity and Chemistry 1. (d) 2. (b) 3. (a) and (d) 4. (d) 5. (a) and (c) Test 15 Chemical Changes and Reactions 1. F, T, F, F, 2. (b) 3. (b) and (f) 4. F, T, T, F 5. Anhydrous; blue 6. (b) 7. (c) 8. T, F, F Test 16 Chemical Changes and Reactions 1. (b) 2. Nitrogen monoxide; carbon monoxide 3. (a) 4. B; D; C 5. (d) 6. Oxidised; reduced 7. T, F, T Test 17 Chemical Changes and Reactions 1. (c) and (d) 2. (a) 3. F, T, T, F, 4. (a) 5. (d) 6. (b) Test 18 Chemical Changes and Reactions 1. 2. 3. 4. (d) (c) F, F, T, T (a) 5. (a), (c) and (d) 6. (c) Test 19 Chemical Changes and Reactions 1. (b) 2. (d) 3. (c) 4. (a) and (c) 5. T, T, T, F 6. (a) 7. (b) and (c) Test 20 Acids, Bases and Salts 1. (a), (c) and (d) 2. F, F, T, 3. (b) and (f) 4. (a) and (d) 5. (b) and (c) 6. (a) 7. (b) and (c) 8. (b) 9. (c) Test 21 Acids, Bases and Salts 1. T, F, T, 2. 7 3. (b), (c) and (d) 4. (d) 5. (d) 6. F, F, T, T, 7. (d) 8. (a) Test 22 Acids, Bases and Salts 1. (a) and (d) 2. (c) 3. (c) 4. Hydrogen; neutral; 7 5. (a) and (b) 6. Step 1 Add nickel powder a little at a time to sulphuric acid until the nickel is in excess. Step 2 Heat the mixture to allow it to react. Step 3 Filter off the excess nickel. Step 4 Boil the filtrate until its volume is reduced to half the original. Step 5 Allow the nickel sulphate to crystallise. 7. (a) 8. (d) Test 23 Acids, Bases and Salts 1. (a) and (d) 2. (b) 3. (d) 4. mineral R calcium carbonate gas T carbon dioxide solid W calcium oxide solid X calcium hydroxide 5. (c) Test 24 Acids, Bases and Salts 1. 2. 3. 4. 5. (c) (b) T, T, F, F (e) (a) and (b) 6. (b) 7. (d) 8. (d) Test 25 Acids, Bases and Salts 1. Protons; some 2. F, T, F, F 3. (a) 4. (b) 5. (b) 6. (b) 7. (d) Test 26 Periodic Table 1. (d) 2. (a), (c) and (d) 3. A transition metal Fe 4. (b) and (c) 5. (c) and (d) 6. (a) 7. (a) and (c) 4. (d) 5. Elements; period 6. (d) 7. T, F, T, F, F 8. (b) 9. Hydrogen Test 27 Periodic Table 1. Bromine passing through sodium iodide Chlorine passing through sodium bromide Iodine passing through sodium chloride 2. (b) Test 28 Periodic Table 1. (c) and (d) 2. (a) and (b) 3. (b) 4. (b) 5. (a) 6. F, F, T, T, 7. (b) 8. (a) Test 29 Metals 1. (c) and (d) 3. (c) 4. Sodium; bauxite 5. (c) 6. (c) Test 30 Metals 1. T, T, F, F, 2. (b) and (d) 3. (c) 4. Sodium; bauxite 5. (c) 6. (c) Test 31 Metals 1. (c) and (d) 2. (a) and (b) 3. (b) 4. (b) 5. (a) 6. F, F, T, T, 7. (b) 8. (a) Test 32 Metals 1. (a) 2. (c) 3. (a) 4. Calcium carbonate 5. Silver; aluminium; tin 6. (d) 7. T, F, F, T 8. (c) Test 33 Metals 1. T, T, F, F, 2. (c) and (d) 3. 12 4. (d) 5. (b) 6. Oxide Test 34 Air and Water 1. 2. 3. 4. 5. 6. (b) (a) (c) Combustion; carbon monoxide (a) (b) and (c) (a) and (e) Test 35 Air and Water 1. T, T, F, F, F, T, 2. T, F, T, F, T, F, 3. (b) 4. (b) 5. (a), (b) and (c) 6. (a), (b) and (c) Test 36 Organic Chemistry 1. (c) 2. (c) 3. Compounds; functional 4. T, T, T, F, 5. (c) Test 37 Organic Chemistry 1. (b) 2. number of single bonds 4 number of double bonds 1 3. (b) 4. T, F, F, T 5. Lighter; lower 6. (c) and (d) 7. (c) Test 38 Organic Chemistry 1. (b) and (d) 2. (a) 3. Hydrogen 4. Paraffin/kerosene; diesel; fuel gas 5. F, T, F, T Test 39 Organic Chemistry 1. (d) 2. (a), (c) and (d) 3. (c) 4. (b), (c) and (e) 5. (b) and (d) Test 40 Organic Chemistry 1. (c) 2. (b) 3. (c), (d), and (e) 4. (c) 5. F, T, F, T, 6. F, F, T, T, 7. (c) and (d) Test 41 Organic Chemistry 1. (a) 2. (d) 3. (a) 4. (c) 5. F, T, T, T, F, 6. (b), (c) and (e) 7. (b) 8. (b) and (c) Test 42 Organic Chemistry 1. (a) 2. (d) 3. (a) 4. (c) 5. F, T, T, T, F, 6. (b), (c) and (e) 7. (b) 8. 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